

Practice for ch6 sec1

1. Changes in energy, color, and odor are indications that a _____ has taken place.
 - a. chemical reaction
 - b. displacement reaction
 - c. precipitation reaction
 - d. separation reaction

2. A description of a chemical reaction that includes the names of all reactants and products, but not the formulas, is called a _____.
 - a. balanced equation
 - b. chemical equation
 - c. skeleton equation
 - d. word equation

3. The substances to the left of the reaction arrow are called _____.
 - a. gases
 - b. precipitates
 - c. products
 - d. reactants

4. The substances to the right of the reaction arrow are called _____.
 - a. products
 - b. reactants
 - c. gases
 - d. precipitates

5. A solid that forms from a solution during a chemical reaction is called a _____.
 - a. precipitate
 - b. substrate
 - c. reactant
 - d. formate

6. What is the missing product in this balanced equation describing the decomposition of water? $2\text{H}_2\text{O}(\text{l}) + \text{energy} \rightarrow \text{O}_2(\text{g}) +$
 - a. $2\text{H}_2(\text{g})$
 - b. $\text{H}_2\text{O}(\text{g})$
 - c. $\text{H}_2(\text{g})$
 - d. energy

7. This type of chemical reaction has energy written as a product.
- endothermic
 - energy-absorbing
 - exothermic
 - light-absorbing
8. A number placed before a formula to indicate how many units of the substance are involved in a chemical reaction is called a _____.
- coefficient
 - substrate
 - subscript
 - superscript
9. What is the coefficient for NaOH in the equation that describes its reaction with CO_2 to produce Na_2CO_3 and H_2O ?
- 1
 - 2
 - 3
 - 4
10. A reaction between magnesium chloride (MgCl_2) and silver nitrate (AgNO_3) yields magnesium nitrate ($\text{Mg}(\text{NO}_3)_2$) and silver chloride (AgCl). What is the coefficient for magnesium nitrate in the balanced equation for the reaction?
- 1
 - 2
 - 3
 - 4
11. This type of chemical reaction has energy written on the right side of a chemical equation.
- endothermic
 - energy-absorbing
 - exothermic
 - light-absorbing