

Worksheet {Sec.1- Chapter 2}

Date: 23/09/2011

Name:

Grade:10

Duration: 20minutes

Section:

I- Multiple Choice:-

1. -----is the smallest particle of matter.

- A. Nucleus.
- B. Atom.
- C. Isotope.
- D. Electron.

2. The mass number of copper is 63. It can be written as:

- A. Copper-63
- B. Cu-62
- C. Cu
- D. None of above.

3. -----equal number of proton plus number of neutrons :-

- A. Atomic mass.
- B. Atomic number.
- C. Mass number.
- D. Average atomic mass.



4. -----What are these numbers?

- A. Atomic mass.
- B. Mass number.
- C. Atomic number.
- D. Average atomic mass.

Silicon 14 Si 28.086	Phosphorus 15 P 30.974	Sulfur 16 S 32.065
Germanium 32 Ge 72.64	Arsenic 33 As 74.922	Selenium 34 Se 78.96

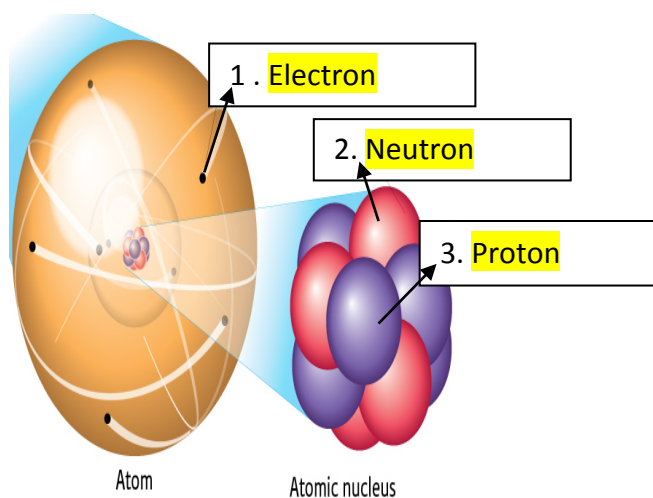


5. -----atoms of the same element, with different number of neutrons:-

- A. Atomic mass.
- B. Atom.
- C. Mass number.
- D. Isotopes.

II- Graph Analysis:-

1. Complete the graph below:-



2. Complete the table :-

Atom	Charge	Mass	Location
Protons	+1	1	Nucleus
Neutron	0	1	Nucleus
Electron	-1	1/1840	Electron cloud

3. Complete the graph with writing the information you can see you from each block in the periodic table :-

5. state of matter		Hydrogen	1. name of element
		1	2. atomic Number
		H	3. symbol
		1.008	4. atomic mass



III- Conceptual Questions:-

1. Why are Atoms Neutral?

The number of protons equal number of electron

2. Why isn't the Atomic Mass a whole number?

Atomic mass is the average of all isotopes of the element .



(THE END)

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Grade: 10-05

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